

## Solubility

Refer to p. 359 - 380 for the following answer.

The formation of most solutions depends on the relative strength of: ( 3 marks)

1)

2)

3)

What is a hydrogen bond? (1 mark)

Explain why water is such a good solvent. (2 marks)

Describe below how ionic compounds dissolve in water: (3 marks)

Describe how molecular compounds dissolve in water: (3 marks)

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J. Kropac

What are conductivity tests used for when dealing with solutions? (1 mark)

If a solution has a high conductivity, what does that suggest about the compounds dissolved in the solution?(2 marks)

How do ion charge effect an ionic compounds solubility? Explain. (3 marks)

How does ion size effect an ionic compounds solubility? Explain. (3 marks)

What is mean by the term 'like dissolves like'? Use a graphic organizer or chart to explain. (2 marks)

How does temperature effect solubility? Use a real world example to explain. (2 marks)

What is a solubility curve and what are they used for? (2 marks)

Use this box to sketch a solubility curve. Be sure to include the solubility of at least three ionic compounds and label both axis. ( 3 marks)

What is pressure and how does it effect solubility? (2 marks)

What is rate of dissolving? ( 1 mark)

List three factors that effect it. (3 marks)

1)

2)

3)

## Quiz Review

You will have a quiz this upcoming **Wednesday Jan 13th**. On the quiz will be the following topics:

i) Solutions

ii) Concentrations

-Practice pg 381 # 41,42

iii) Solution Stoichiometry

-Practice pg 417 # 18

iv) Dilutions

-Practice pg 386 # 51

v) Acids and Bases

- List some common characteristics of acids and bases
- How did Arrhenius define an acid and base
- How did Bronsted-Lowry describe an acid and base
- What is pH scale

vi) Acid and Base Reactions

- What are the four types of reactions acids undergo
- What is titration used to determine