

Galvanic Cells

Some things to consider when drawing galvanic cell:

- 1) Electrons always flow from the anode to the cathode.
- 2) The anode always undergoes oxidation.
- 3) The metal that is higher in the activity series is placed at the anode.

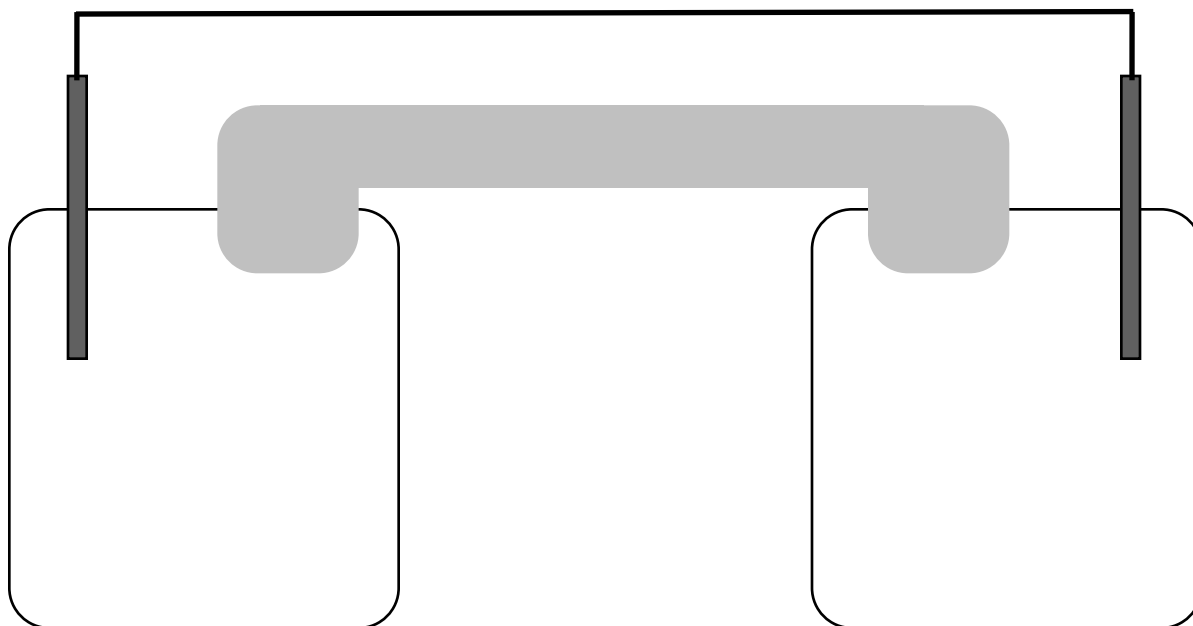
Galvanic Cell Worksheet

Consider a galvanic cell that contains nickel and copper and a salt bridge containing sodium nitrate, NaNO_3

- 1) Start by writing out your half reactions:

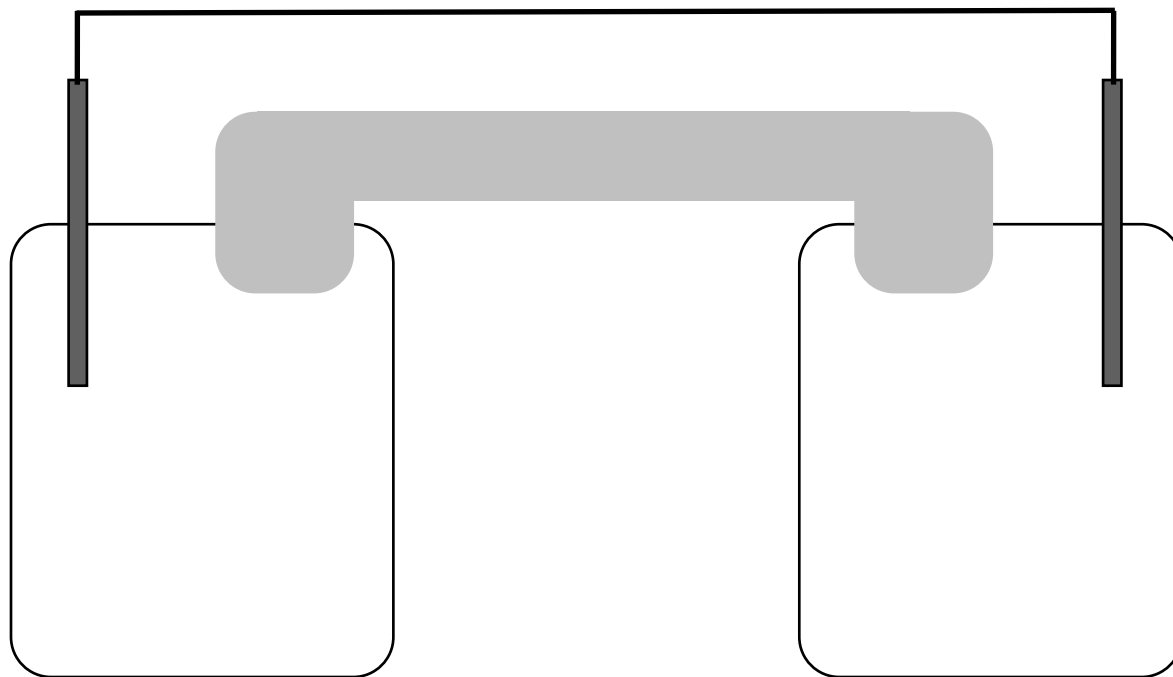
Hint: the metal that is higher in the activity series will undergo oxidation.

- 2) Now label both the anode and the cathode
- 3) Now show the flow of electrons.
- 4) Show the flow of ions from the salt bridge.

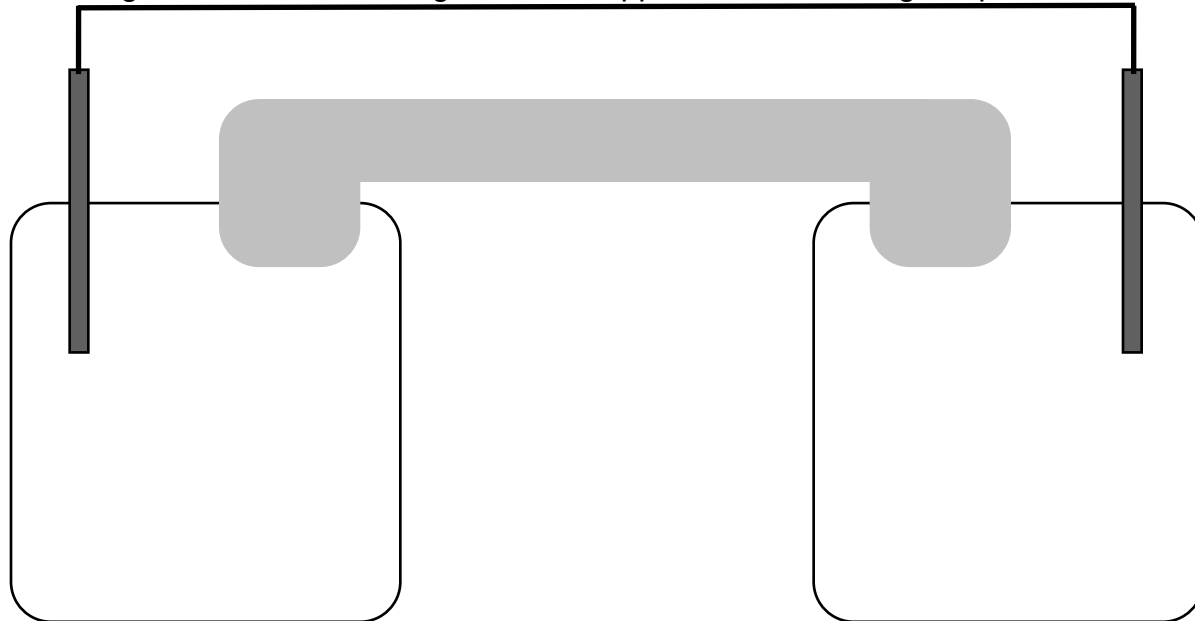


SCH 4C
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Draw a galvanic cell containing platinum and iron (II) and a salt bridge of potassium hyponitrite, KNO_2 .

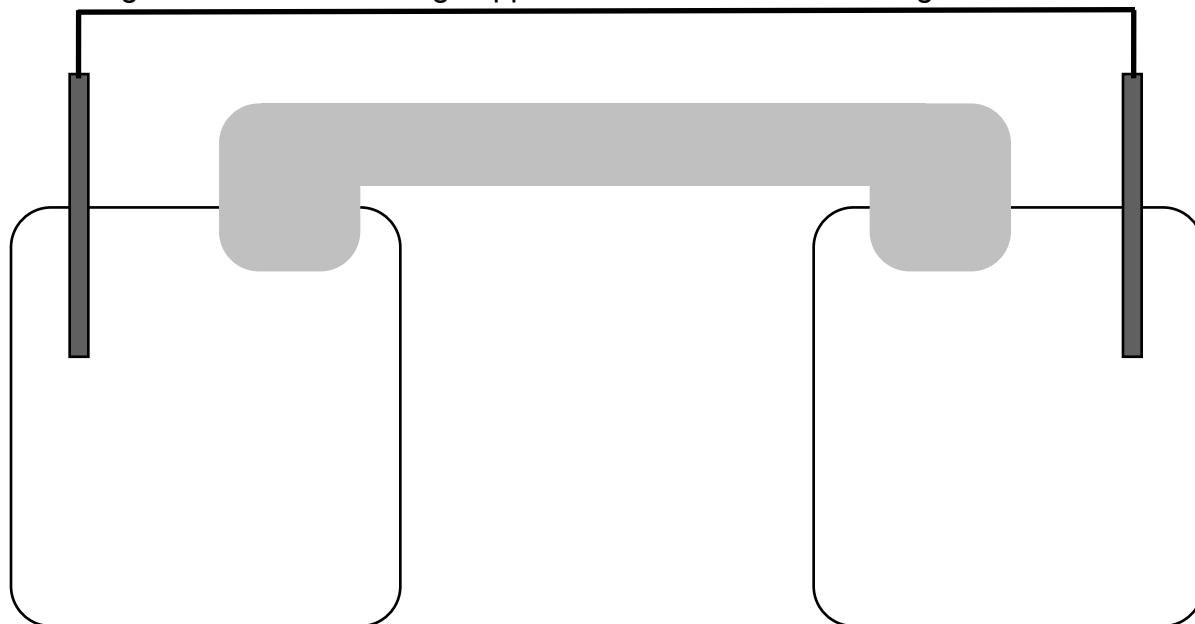


Draw a galvanic cell containing zinc and copper with a salt bridge of potassium chloride.



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Draw a galvanic cell containing copper and silver with a salt bridge of sodium chloride.



Draw a galvanic cell containing zinc and lead with a salt bridge of lithium chloride.

