

**Assignment One: Binary Compounds**

Write the IUPAC name for each of the following binary compounds:

LiF \_\_\_\_\_

SrI<sub>2</sub> \_\_\_\_\_

CaO \_\_\_\_\_

SO<sub>2</sub> \_\_\_\_\_

Al<sub>2</sub>O<sub>3</sub> \_\_\_\_\_

CF<sub>4</sub> \_\_\_\_\_

BF<sub>3</sub> \_\_\_\_\_

NaBr \_\_\_\_\_

PCl<sub>5</sub> \_\_\_\_\_

ZnS \_\_\_\_\_

CuBr<sub>2</sub> \_\_\_\_\_

CoI<sub>3</sub> \_\_\_\_\_

Ag<sub>3</sub>P \_\_\_\_\_

CO \_\_\_\_\_

Sb<sub>2</sub>S<sub>5</sub> \_\_\_\_\_

P<sub>4</sub>O<sub>10</sub> \_\_\_\_\_

NaF \_\_\_\_\_

MnS \_\_\_\_\_

AuCl<sub>3</sub> \_\_\_\_\_

HgO \_\_\_\_\_

CuCl \_\_\_\_\_

NiF<sub>2</sub> \_\_\_\_\_

CBr<sub>4</sub> \_\_\_\_\_

KI \_\_\_\_\_

SF<sub>6</sub> \_\_\_\_\_

Write the chemical formula for each of the following binary compounds:

iron (III) sulfide \_\_\_\_\_

calcium nitride \_\_\_\_\_

rubidium bromide \_\_\_\_\_

barium oxide \_\_\_\_\_

chromium (II) chloride \_\_\_\_\_

silver sulfide \_\_\_\_\_

bismuth (III) fluoride \_\_\_\_\_

gold (III) oxide \_\_\_\_\_

bismuth (V) chloride \_\_\_\_\_

aluminum bromide \_\_\_\_\_

copper (I) hydride \_\_\_\_\_

sulfur trioxide \_\_\_\_\_

francium sulfide \_\_\_\_\_

antimony (III) oxide \_\_\_\_\_

carbon tetrachloride \_\_\_\_\_

iron (III) oxide \_\_\_\_\_

cesium phosphide \_\_\_\_\_

boron trichloride \_\_\_\_\_

calcium sulfide \_\_\_\_\_

zinc chloride \_\_\_\_\_

lead (IV) bromide \_\_\_\_\_

dinitrogen pentoxide \_\_\_\_\_

cobalt (III) iodide \_\_\_\_\_

potassium hydride \_\_\_\_\_

barium sulfide \_\_\_\_\_

nitrogen trichloride \_\_\_\_\_