


Two Activity Series			
Metals	Decreasing Activity	Halogens	
lithium		fluorine	
potassium		chlorine	
calcium		bromine	
sodium		iodine	
magnesium			
aluminum			
zinc			
chromium			
iron			
nickel			
tin			
lead			
HYDROGEN*			
copper			
mercury			
silver			
platinum			
gold			

- In your own words what does 'Activity Series' mean?
- For each of the following, use the activity series to determine which single displacement reactions will proceed. For the reactions that do occur, predict the products and complete and balance the equation. Note reactions that do not occur with NR.
 - $\text{Zn}_{(s)} + \text{CuCl}_{2(aq)} \rightarrow$
 - $\text{Br}_{(s)} + \text{CaCl}_{2(aq)} \rightarrow$
 - $\text{Pb}_{(s)} + 2\text{HCl}_{(aq)} \rightarrow$
 - $\text{I}_{2(s)} + \text{HCl}_{(aq)} \rightarrow$
 - $\text{Au}_{(s)} + \text{ZnSO}_{4(aq)} \rightarrow$
 - $\text{Sn}_{(s)} + 2\text{AgNO}_{3(aq)} \rightarrow$
 - $2\text{Al}_{(s)} + 3\text{H}_2\text{O}_{(l)} \rightarrow$
 - $\text{Al}_{(s)} + \text{ZnSO}_{4(aq)} \rightarrow$