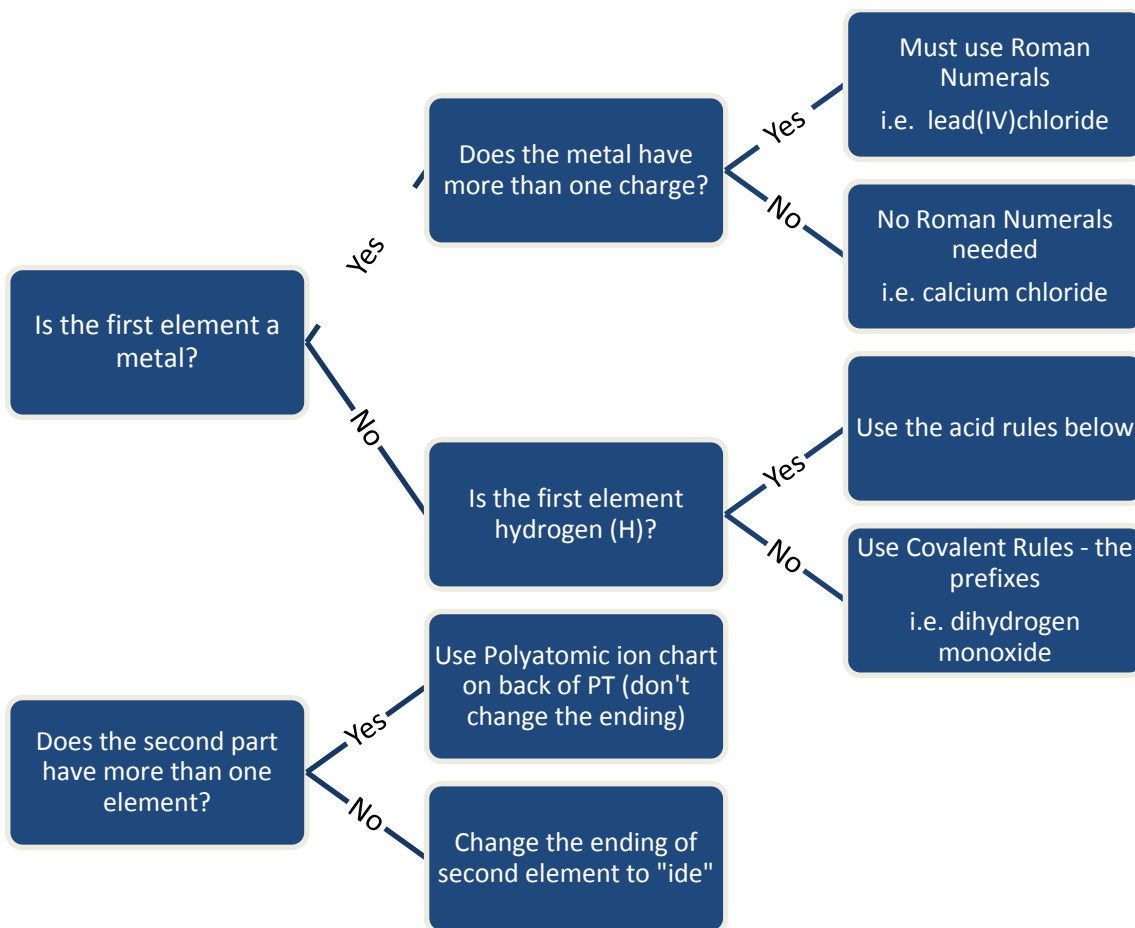


Naming Summary | SCH3U Chemistry



Acid Rules:

Hydrogen and one other element:

hydro_____ic acid – formed from the “ide” ion and hydrogen
 hydrochloric acid HCl hydrosulfuric acid H₂S

Hydrogen and two or more elements

_____ic acid – formed from the “ate” ion and hydrogen
 Chloric acid HClO₃ sulfuric acid H₂SO₄

_____ous acid – formed from the “ite” ion and hydrogen
 Chlorous acid sulfurous acid H₂SO₃

Remembering your Polyatomic Ions:

Memorize the “ate” ion

“-ate” + 1 Oxygen	“per___ate”	SO ₅ ²⁻	persulfate
“-ate”	“-ate”	SO ₄ ²⁻	sulfate
“-ate” – 1 Oxygen	“-ite”	SO ₃ ²⁻	sulfite
“-ate” -2 Oxygen	“hypo___ite”	SO ₂ ²⁻	hyposulfite

Prefixes – COVALENT ONLY!

1 – mono	2 – di	3 – tri	4 – tetra	5 – penta
6 – hexa	7 – hepta	8 – octa	9 – nona	10 – deca

Hydrates:

Hydrates – have water attached to them i.e. CuSO₄•5H₂O

Name the ionic compound – copper(II)sulfate

Then add the number of “hydrates” so pentahydrate

Name - copper(II)sulfate pentahydrate